



13th International Conference on Catalysis in Membrane Reactors



Preparation and hydrogen permeation studies of ultrathin palladium and carbon membranes

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NORDIO², MARTIN VAN SINT ANNALAND² FAUSTO GALLUCCI².

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July 10-13 2017. Houston, Texas

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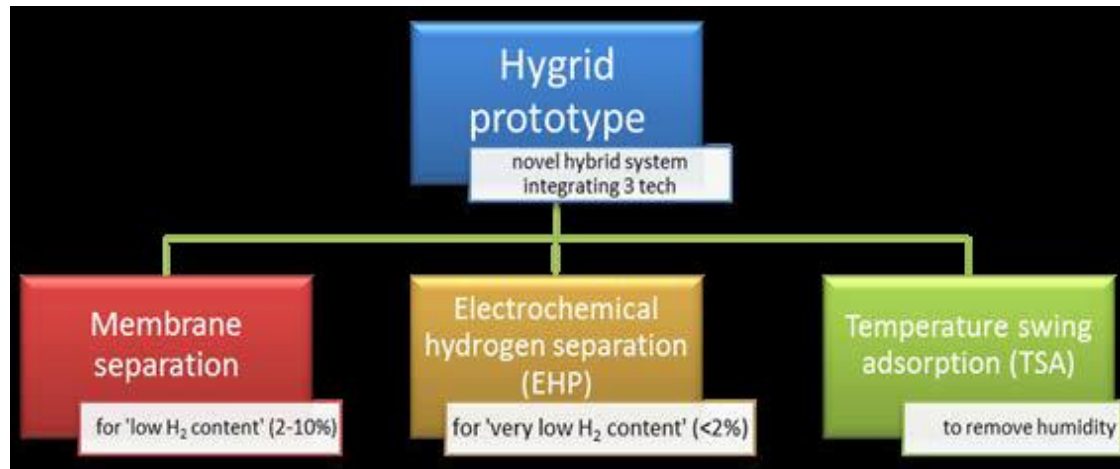


HyGrid

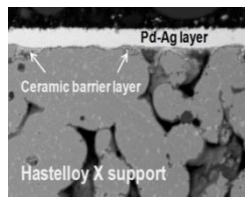


Objective

Developing of an advanced high performance, cost effective separation technology for direct separation of hydrogen from natural gas networks



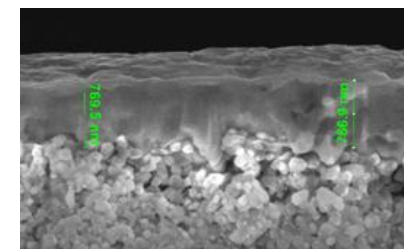
Thin



(3-5 μm thick)



Ultra thin



(≈ 1 μm thick)

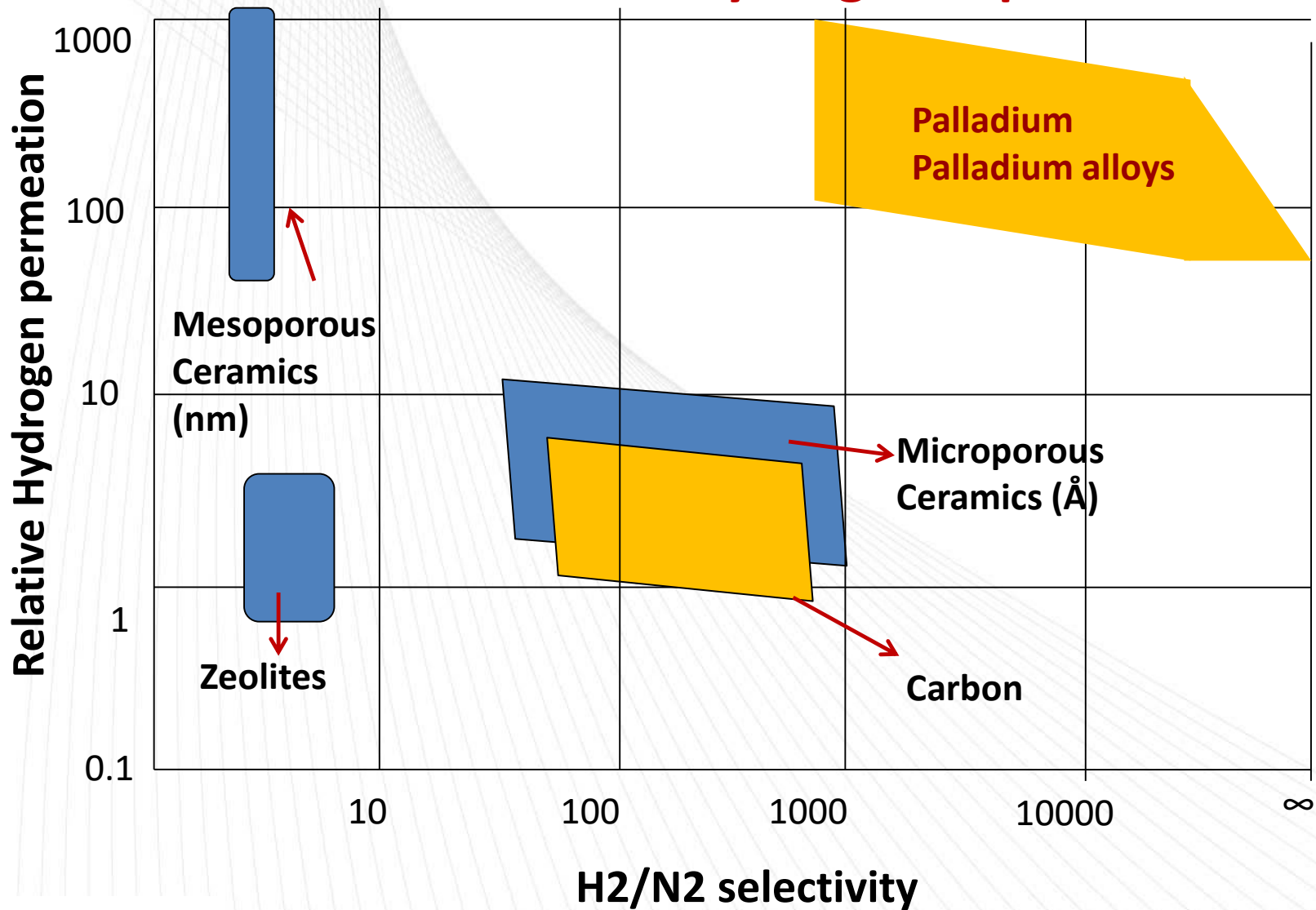
Thin and ultra thin supported Pd/Ag membranes

D. ALFREDO PACHECO TANAKA¹, JON MELENDEZ¹,
EKAIN FERNANDEZ¹, MARIA NORDIO²,
FAUSTO GALLUCCI²

Poster

ULTRA-THIN PALLADIUM-SILVER MEMBRANES FOR PURE
HYDROGEN PRODUCTION AND SEPARATION: MODELLING
AND EFFECT OF SWEEP GAS

Membranes for hydrogen separation

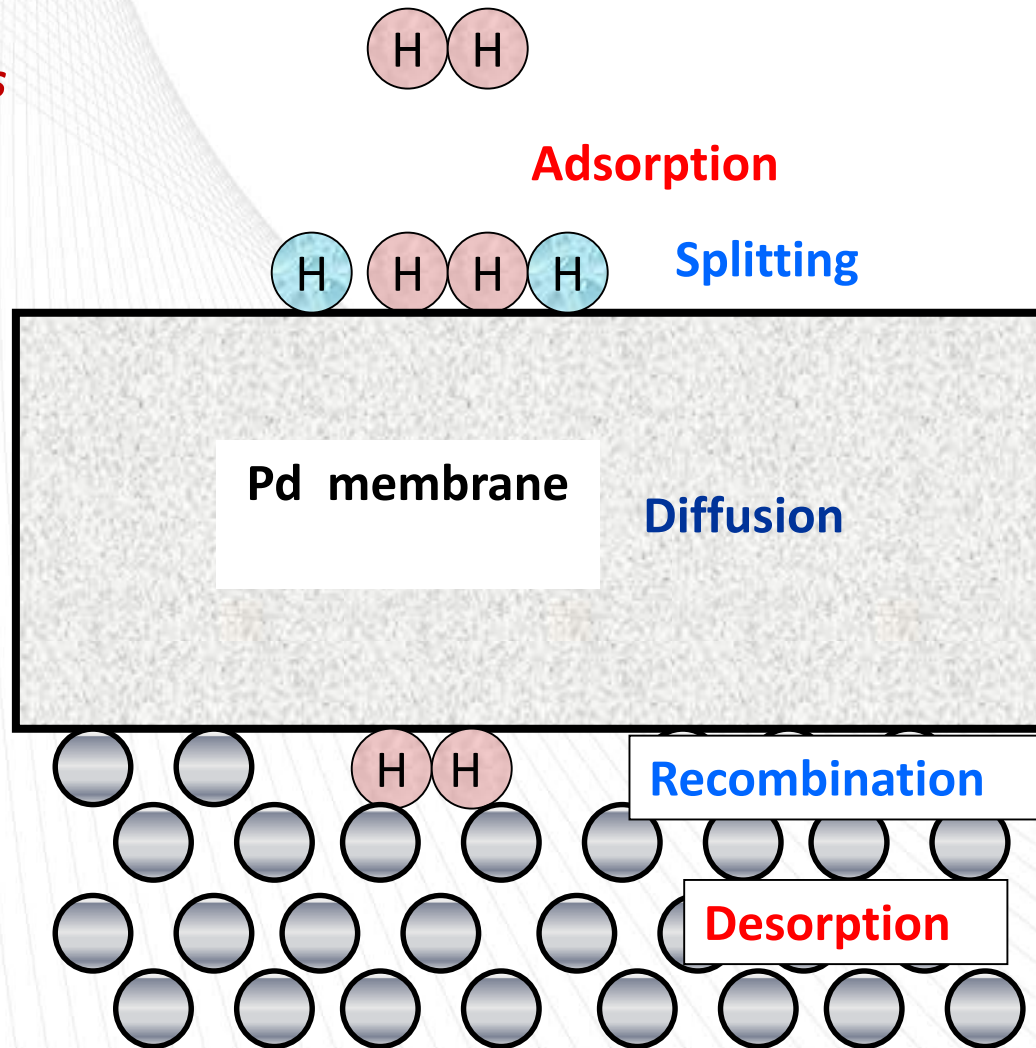


H₂ permeation in Pd membranes

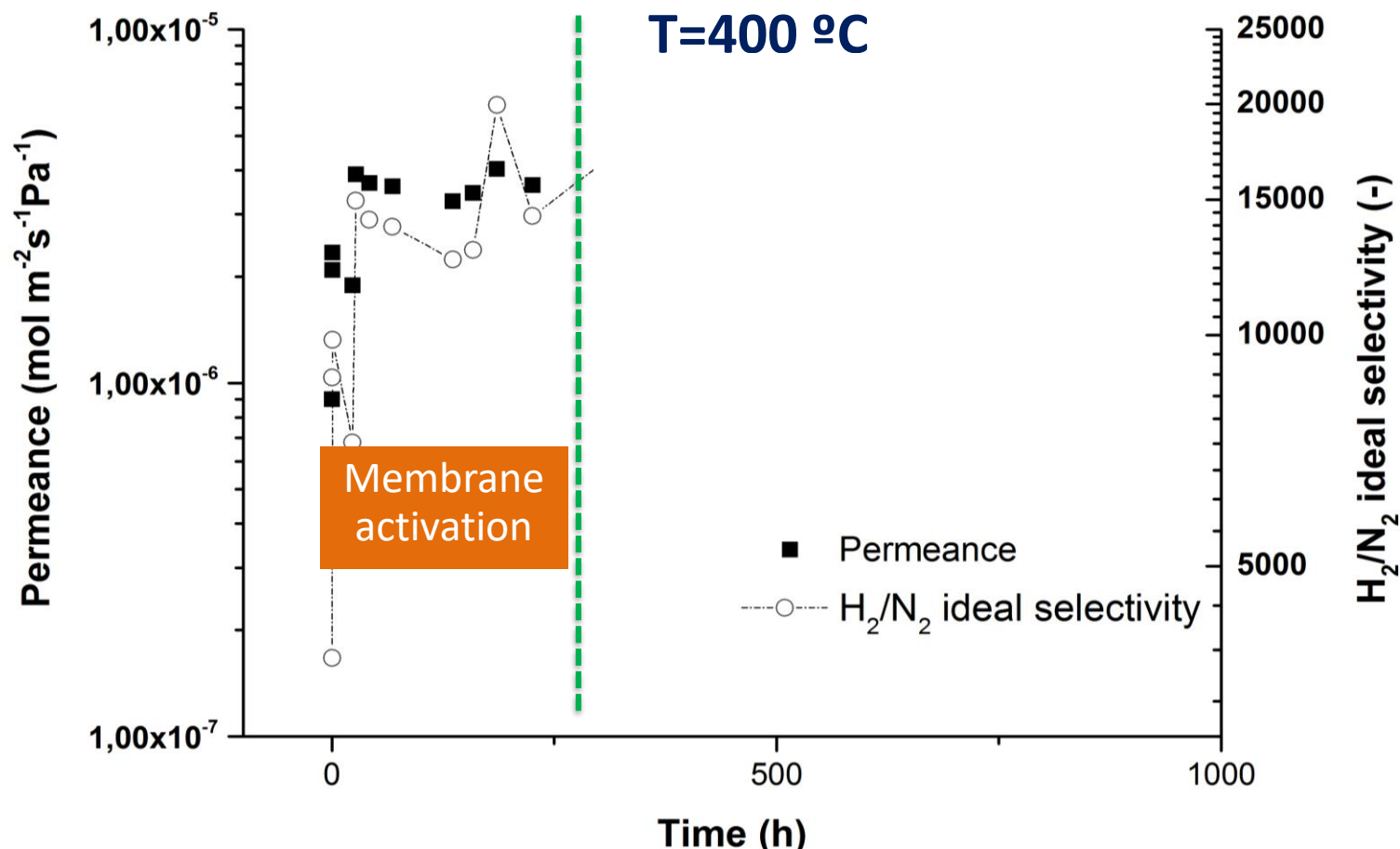
Ultra-thin membranes

*-Surface interference
- H₂ splitting*

resistance of support



Ceramic supported Pd-based (3-4 μm thick) membrane

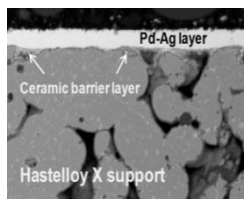


Helmi et al., *Molecules* 21 (2016) 376

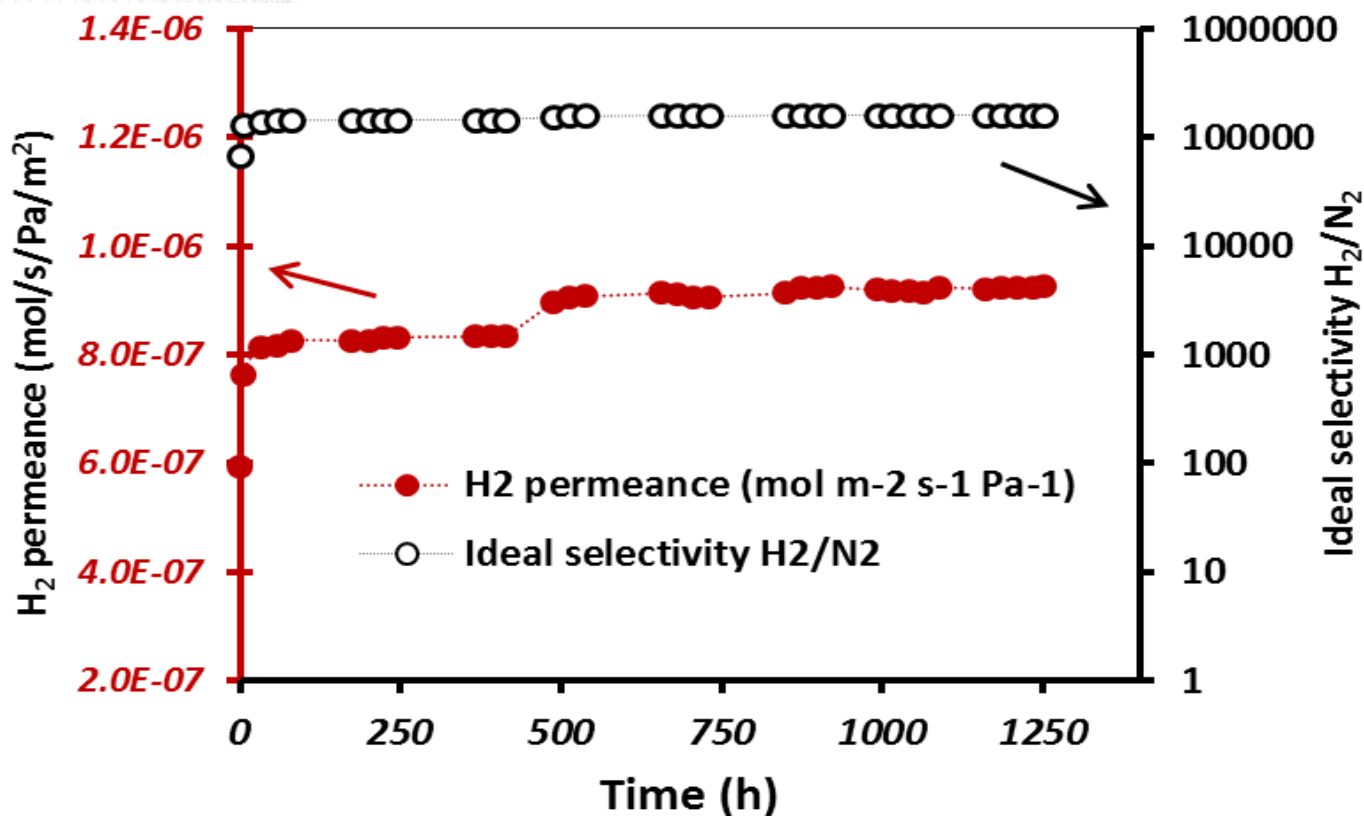
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Metallic supported Pd-Ag membranes

Long term permeation test at **400 °C**



(5 μm thick)

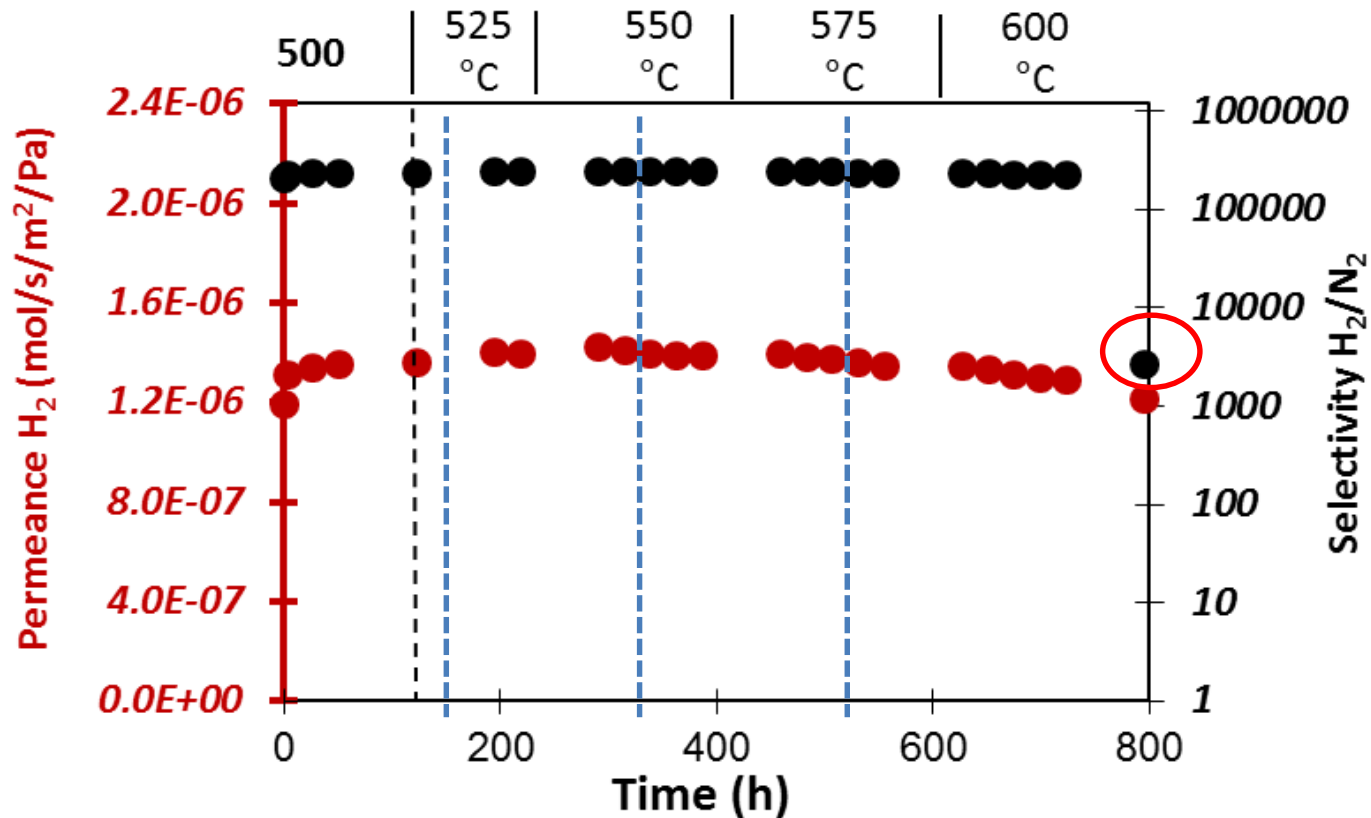
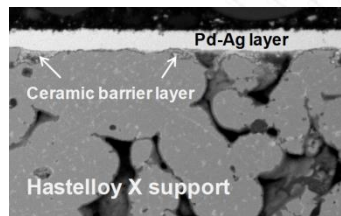


- H₂ permeance 400 °C: **$0.9 \times 10^{-6} \text{ mol m}^{-2} \text{ s}^{-1} \text{ Pa}^{-1}$**
- H₂/N₂ ideal selectivity 400 °C: **>150,000 for >1200 h**

Medrano et al., *Int. J. Hydrogen Energy* 41 (2016) 8706-8718

Metallic supported Pd-Ag membranes

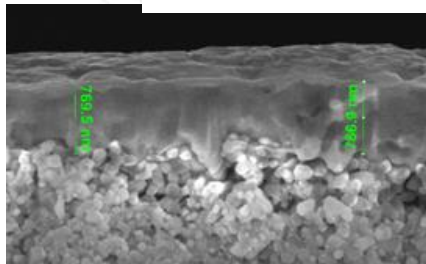
Long term permeation test at more than 500 °C



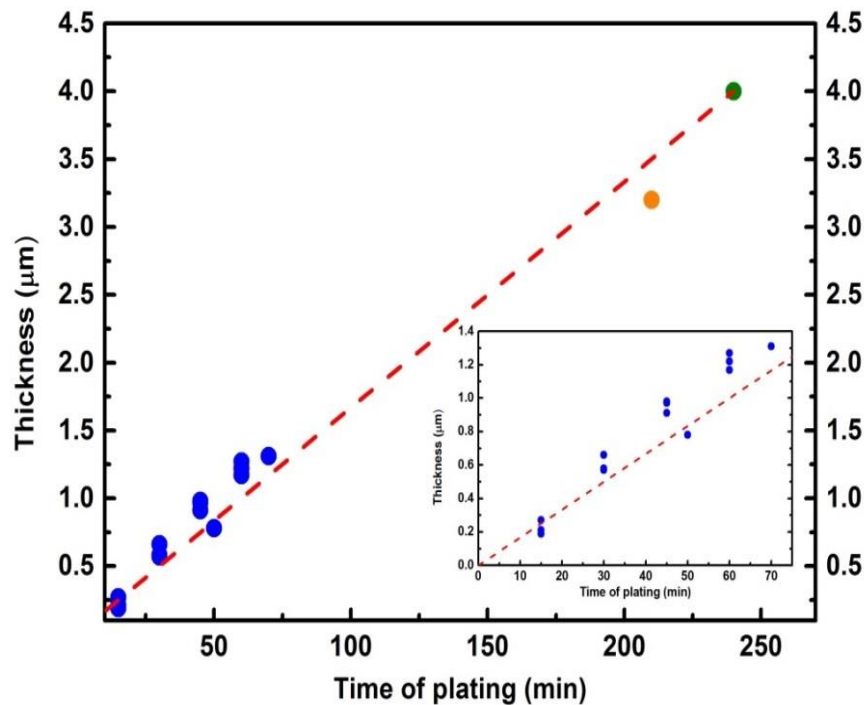
- At >500 °C and after >800 h, H₂/N₂ ideal selectivity dropped to from 223000 to **2650**

Medrano et al., *Int. J. Hydrogen Energy* 41 (2016) 8706-8718

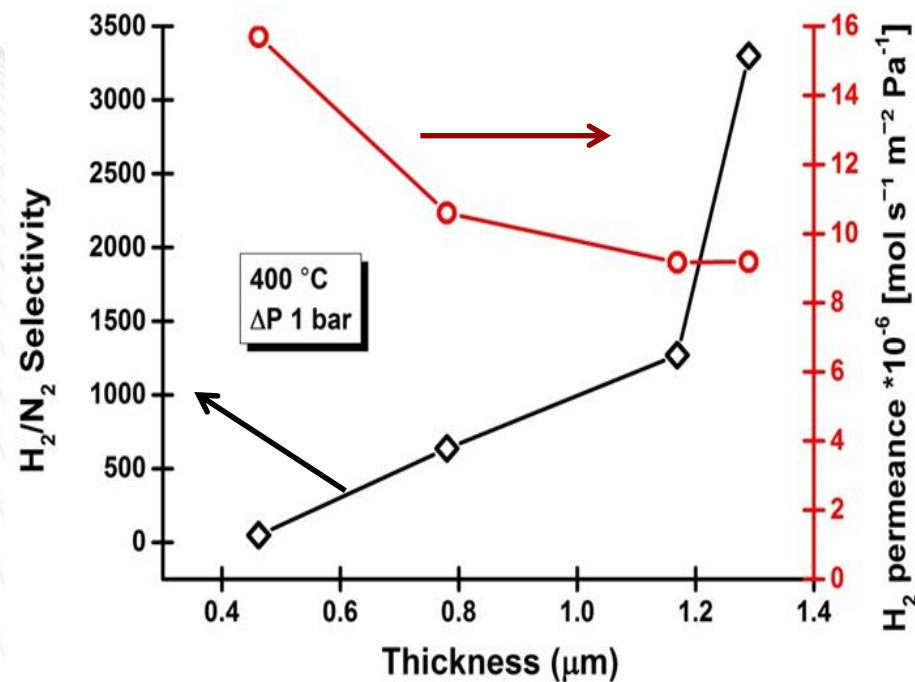
Ultra-thin Pd-Ag membranes $\approx 1\mu\text{m}$



Thickness vs plating time

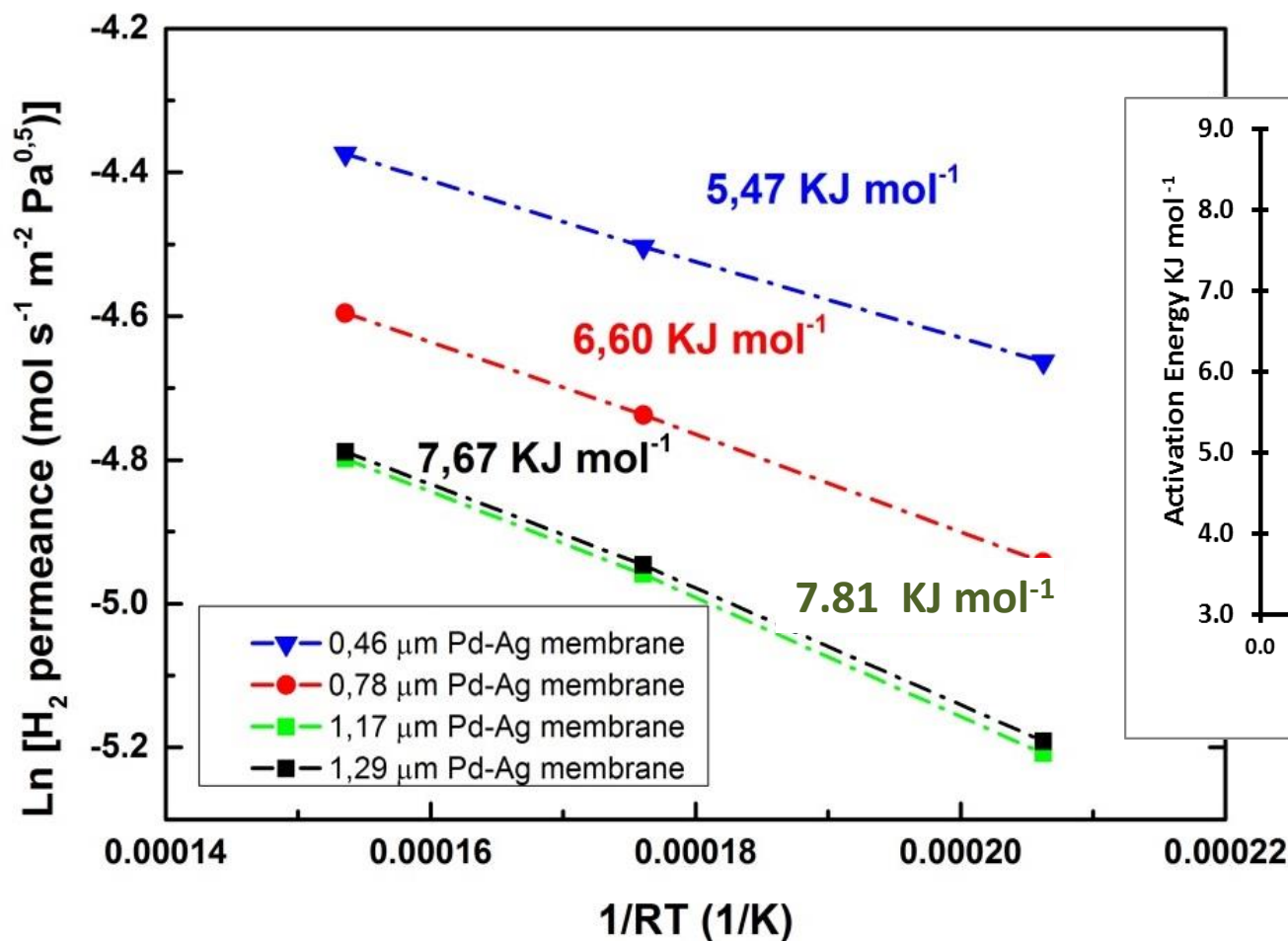


H₂/N₂ selectivity and H₂ permeance vs thickness

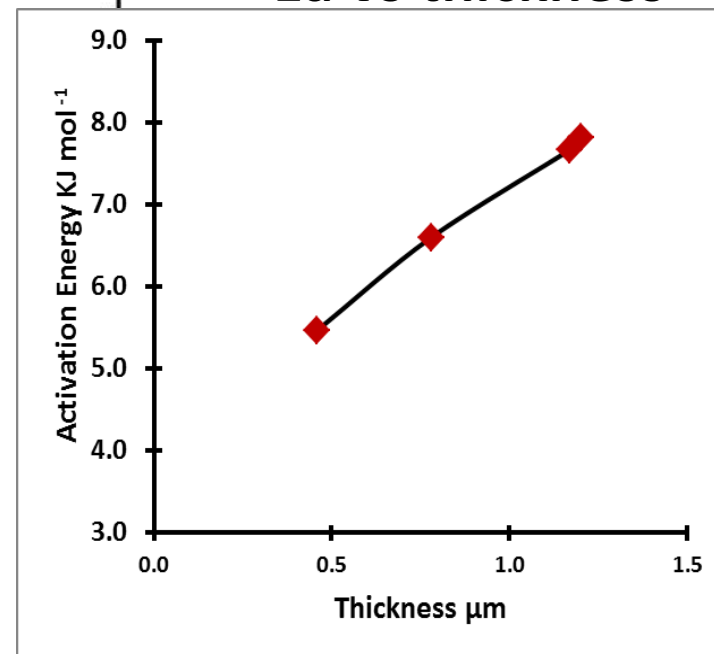


Melendez et al., *Journal of Membr. Sci.* 528 (2017) 12–23

Activation energy of Ultra-thin Pd-Ag membranes



Ea vs thickness

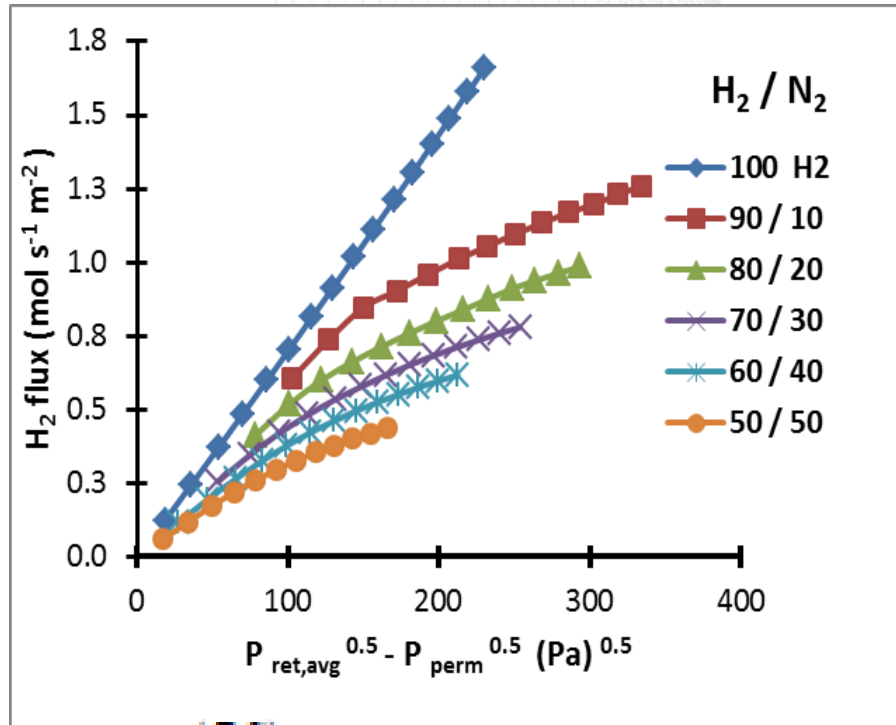


Melendez et al., *Journal of Membr. Sci.* 528 (2017) 12–23

H₂ flux at various gas mixtures and pressures at 400°C

Membrane thickness 1.30 μm

Effect of dilution H₂ and N₂

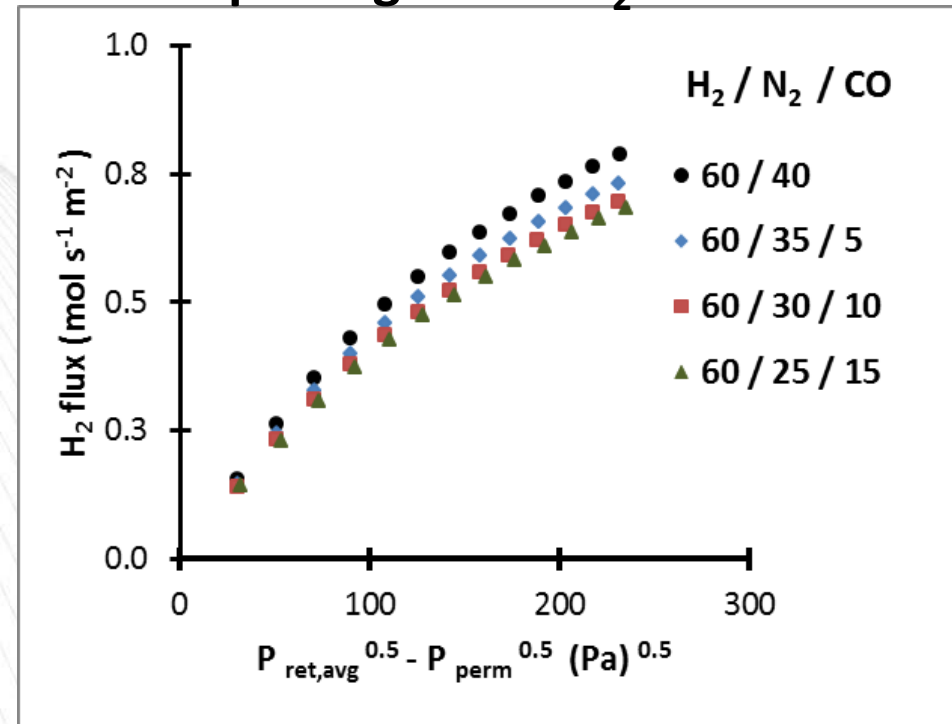


$$J = \left(\frac{DS}{l} \right) e^{-(E_a/RT)} (P_0^n - P_1^n)$$

Effect of CO

At H₂ / N₂ 60/40

Replacing some N₂ with CO

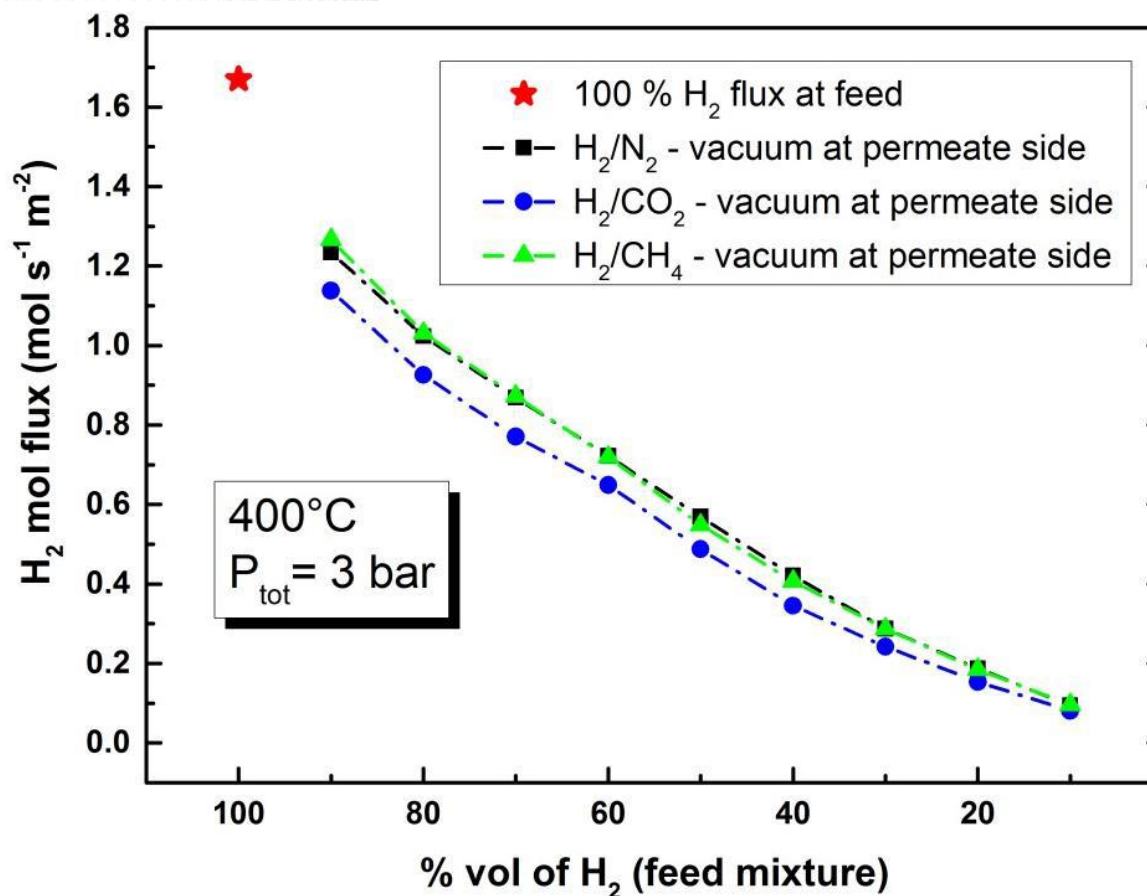


Melendez et al., Journal of Membr. Sci. 528 (2017) 12–23

H₂ flux at various gas mixtures and pressures at 400°C

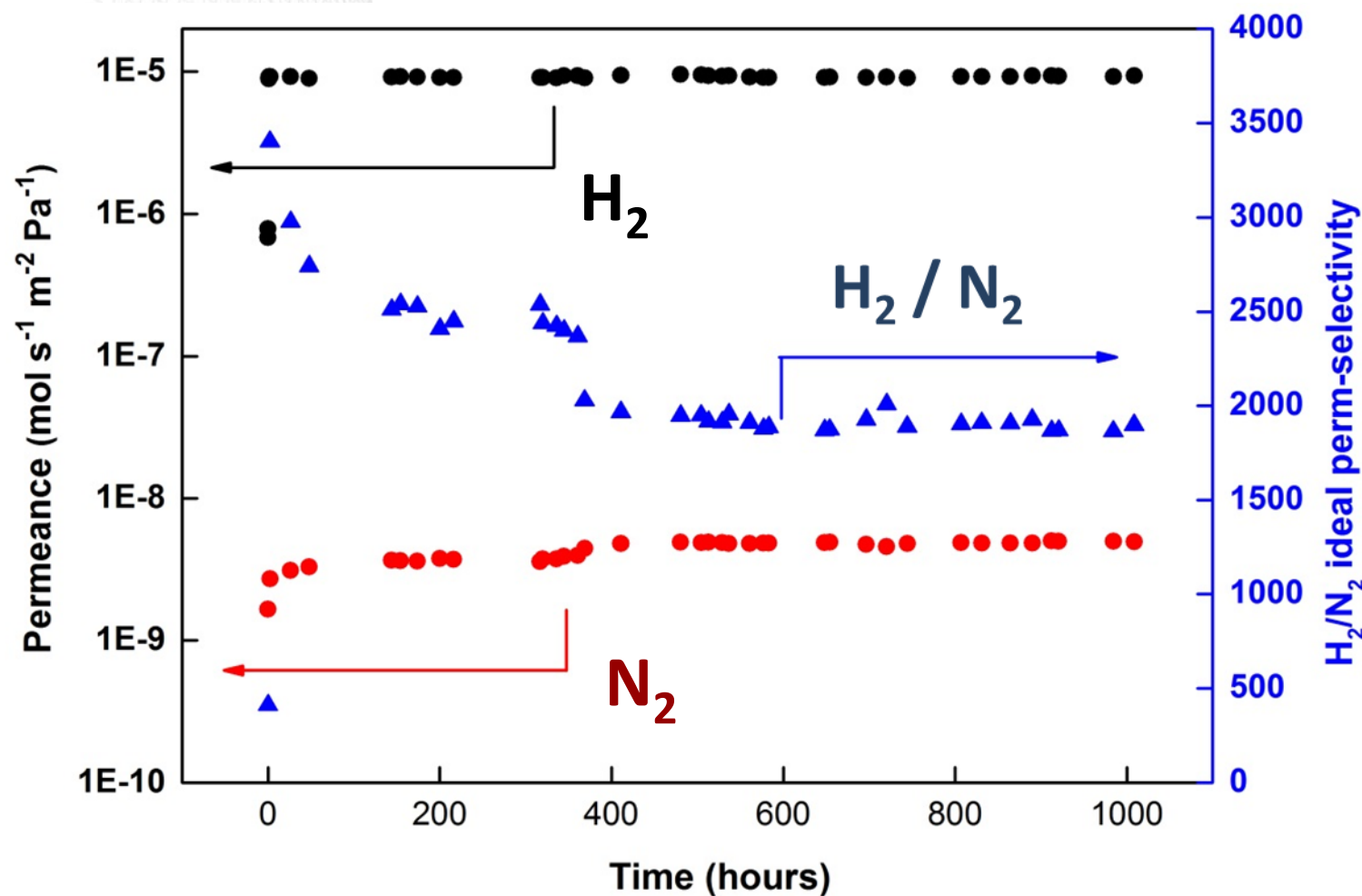
Thickness 1.30 μm

3 bar total pressure, vacuum at the permeate



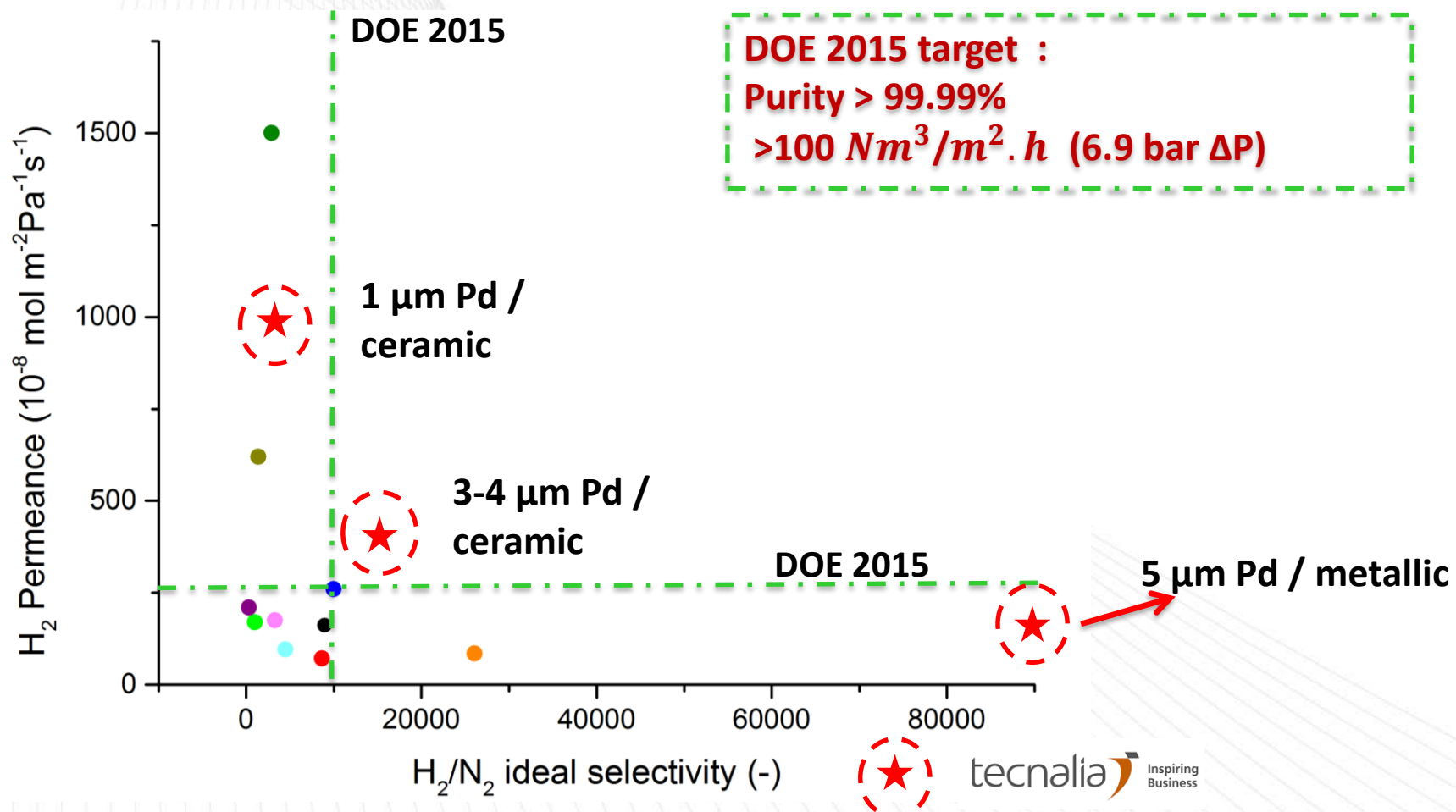
Melendez et al., *Journal of Membr. Sci.* 528 (2017) 12–23

Long term test Pd-Ag membranes $\approx 1,3 \mu\text{m}$



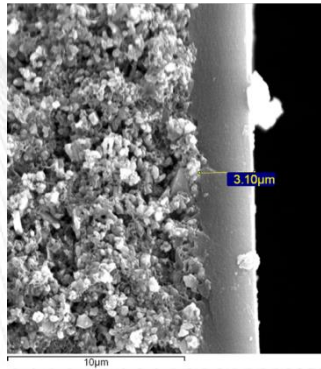
Melendez et al., *Journal of Membr. Sci.* 528 (2017) 12–23

Pd alloy membrane properties at TECNALIA



Gallucci et al., Chem. Eng. Sci 92 (2013) 40–66

Carbon Molecular Sieves Membranes (CMSM)



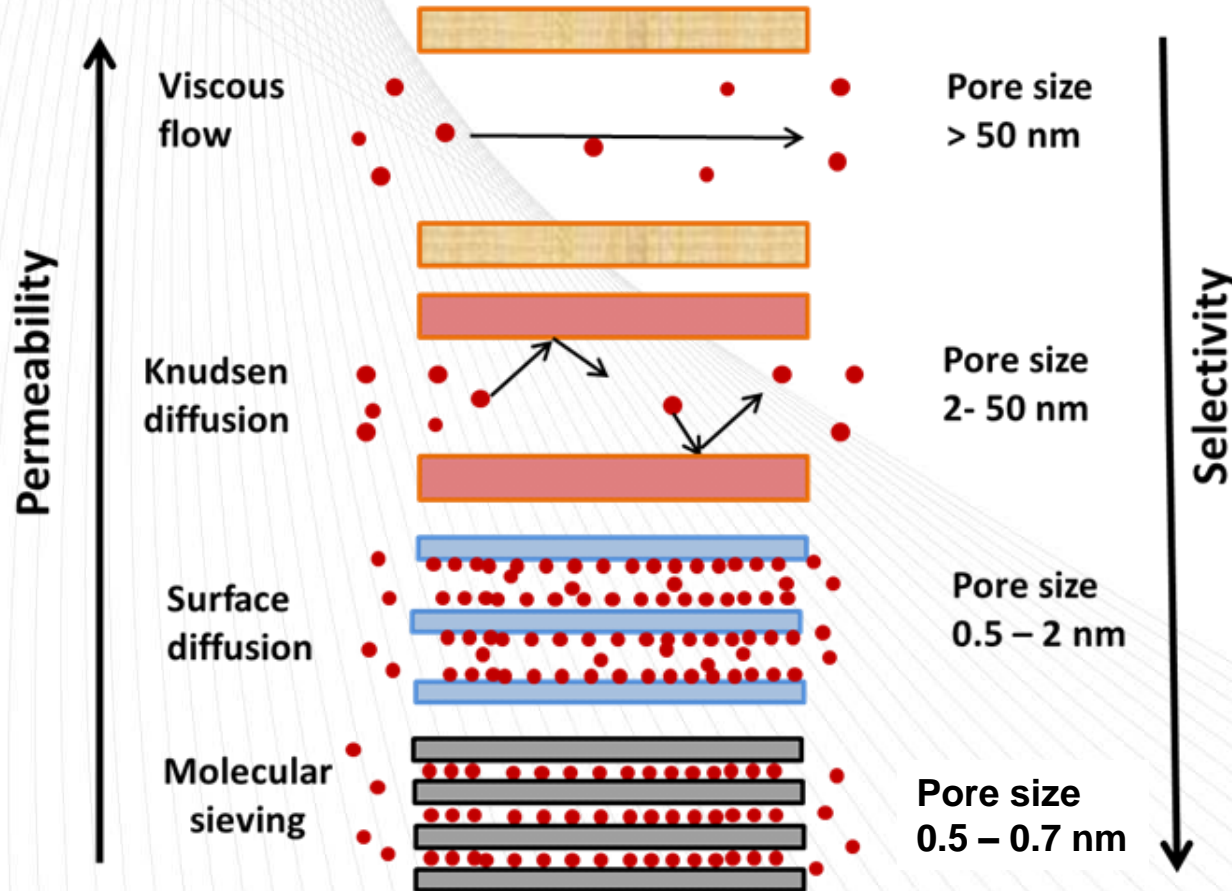
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Poster

**BIOGAS UPGRADING WITH CARBON MEMBRANES:
EXPERIMENTAL DEMONSTRATION AND BENEFITS
EVALUATION**

Porous membranes for gas separation

Mechanism



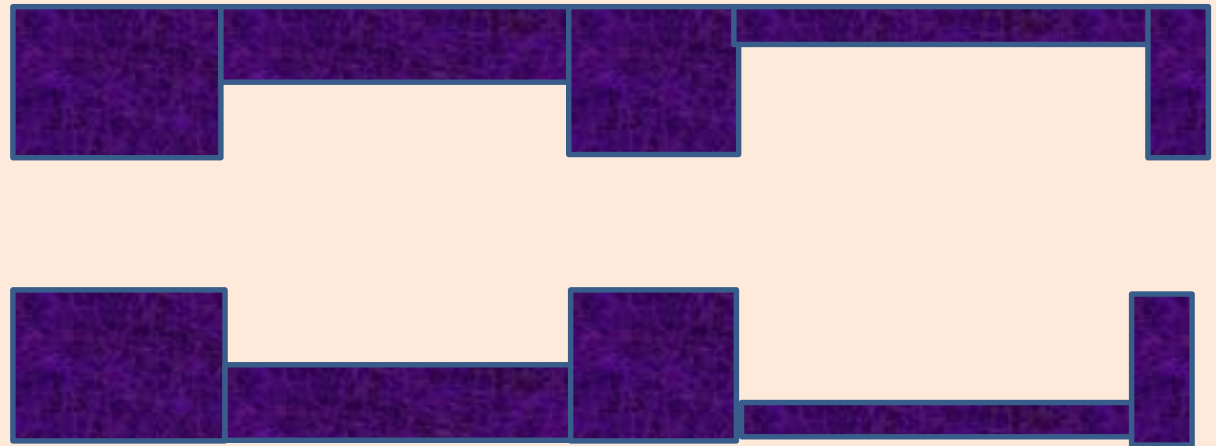
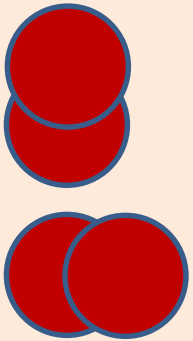
Molecular sieves membranes

Zeolite membranes

- Difficult to obtain crack-free
- Intercrystallite voids affecting selectivity

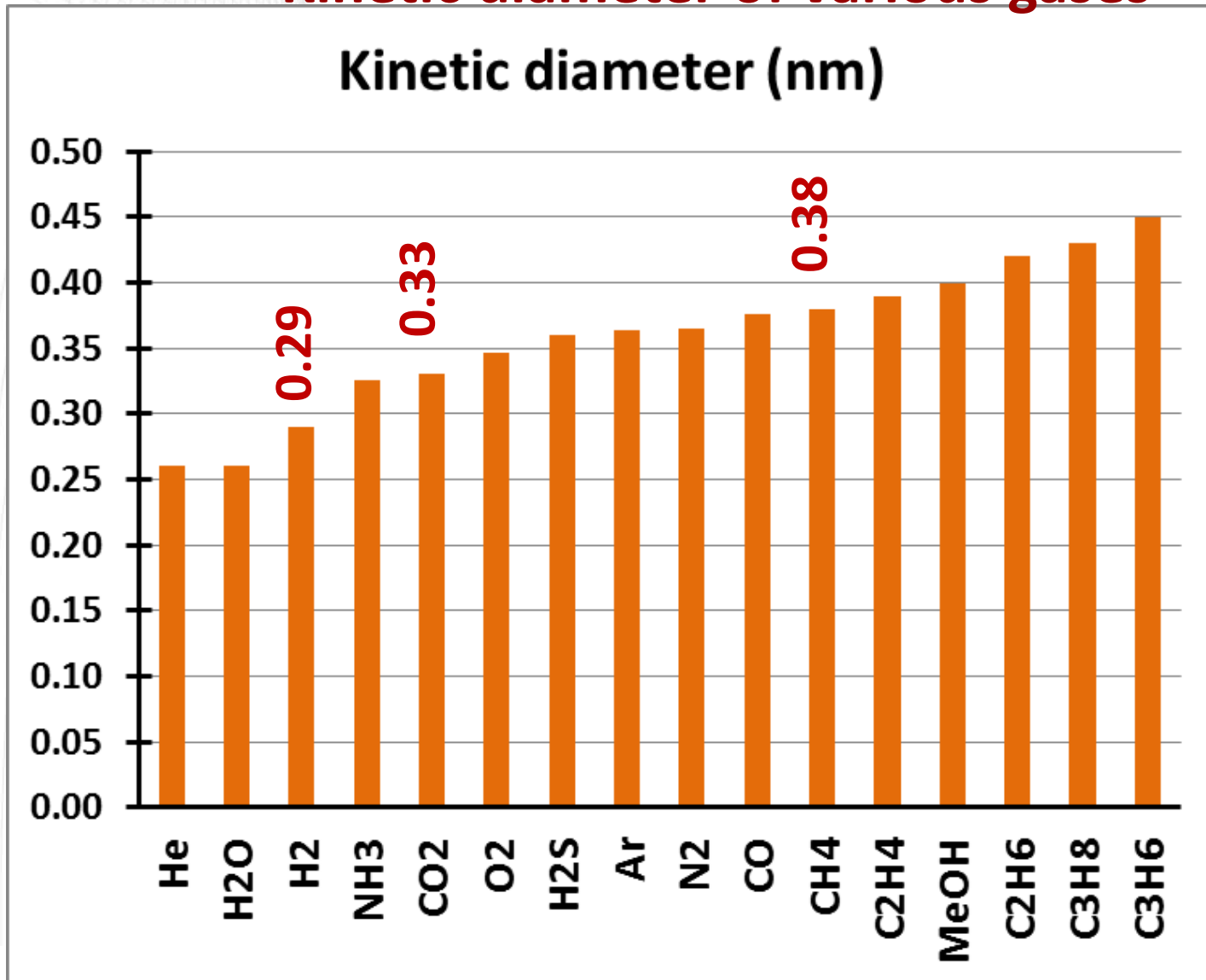
Carbon Molecular sieves membranes

Carbonization of a polymer precursor under inert atmosphere or vacuum

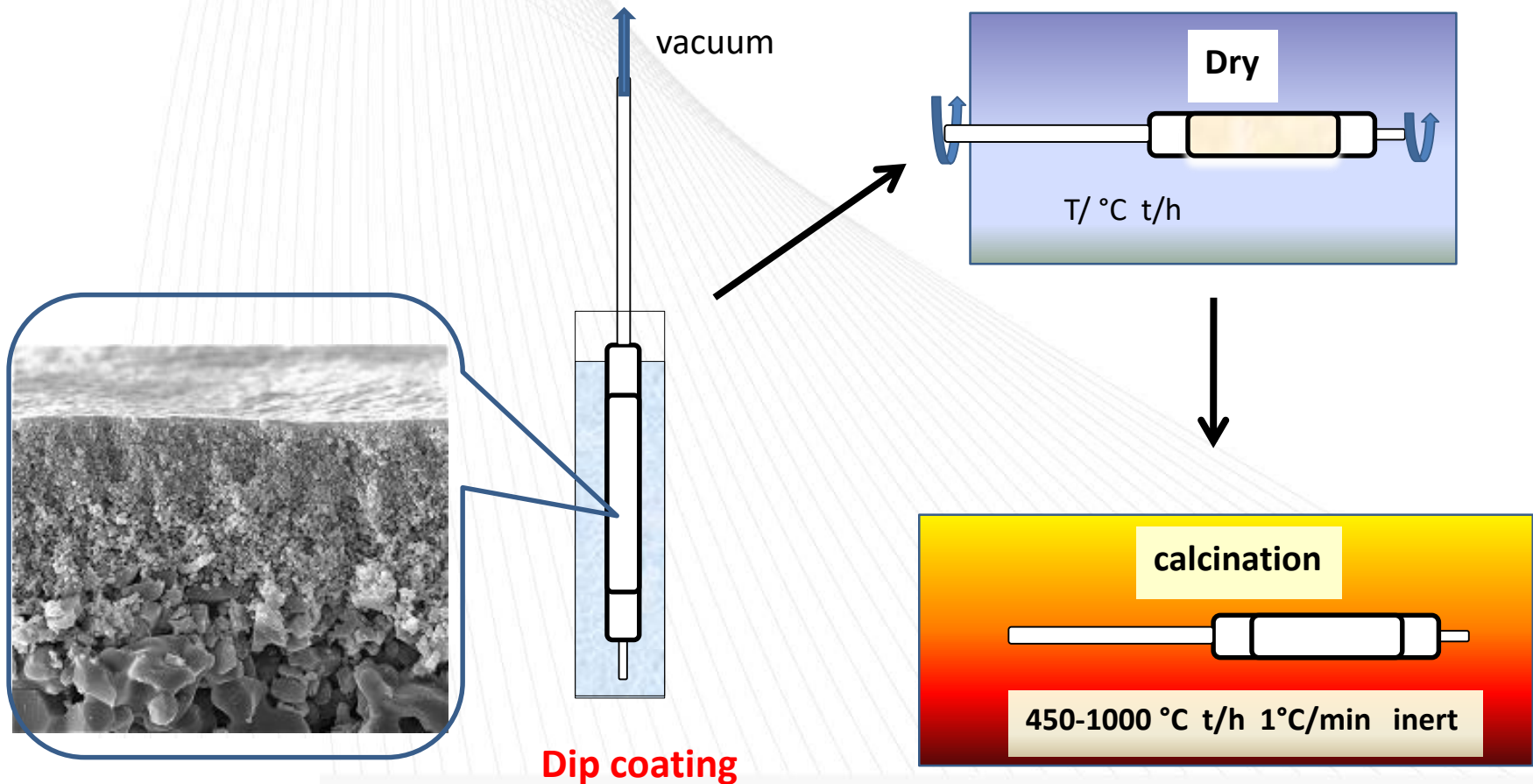


Presence of micropores and ultramicropores

Kinetic diameter of various gases

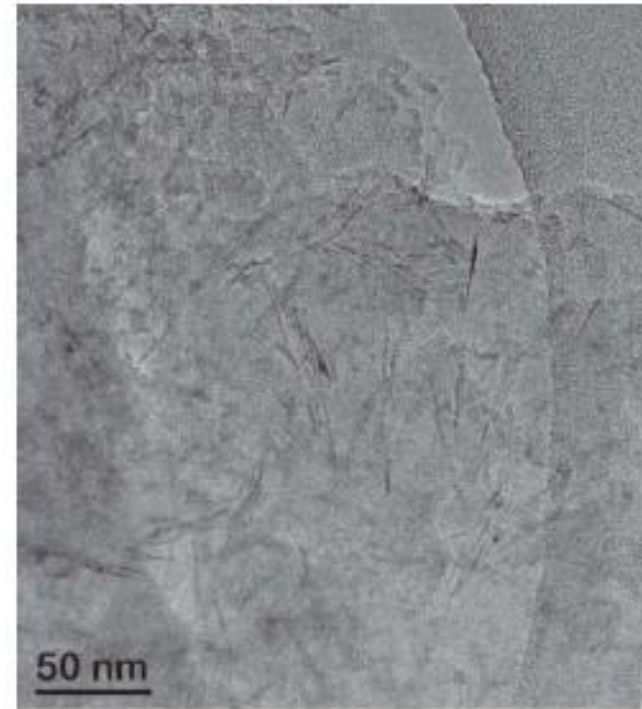
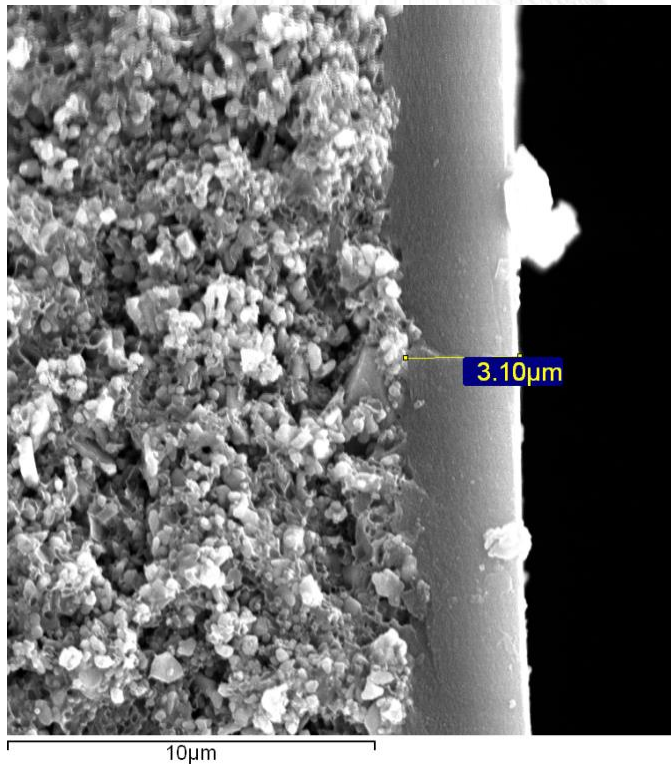


Preparation of composite-Alumina-CMSM



Dip coating

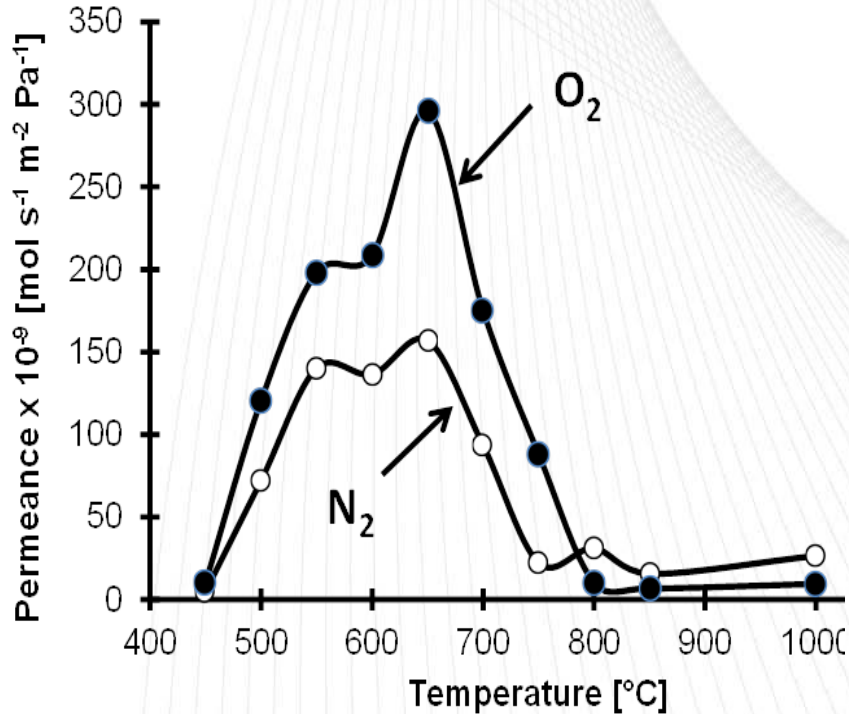
SEM and TEM of the Al-CMSM



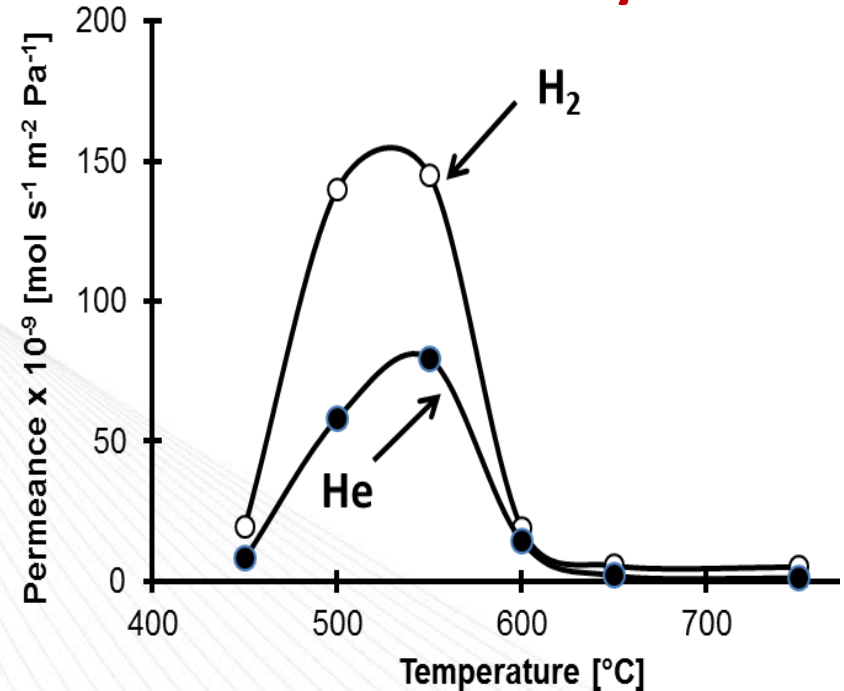
Composite phenolic resin CMSM, exeira, Campo, Tanaka, Llosa, Magen, Mendes, Carbon 49 (2011) 4348

Effect of carbonization temperature

Fresh membrane



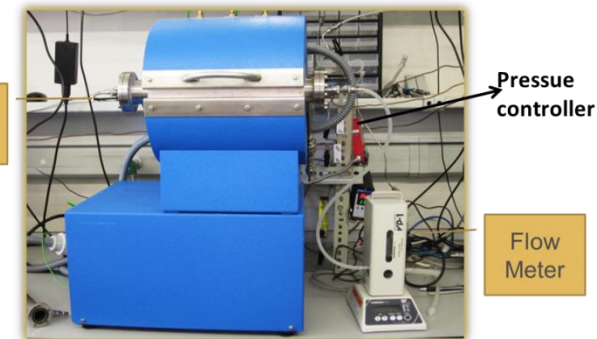
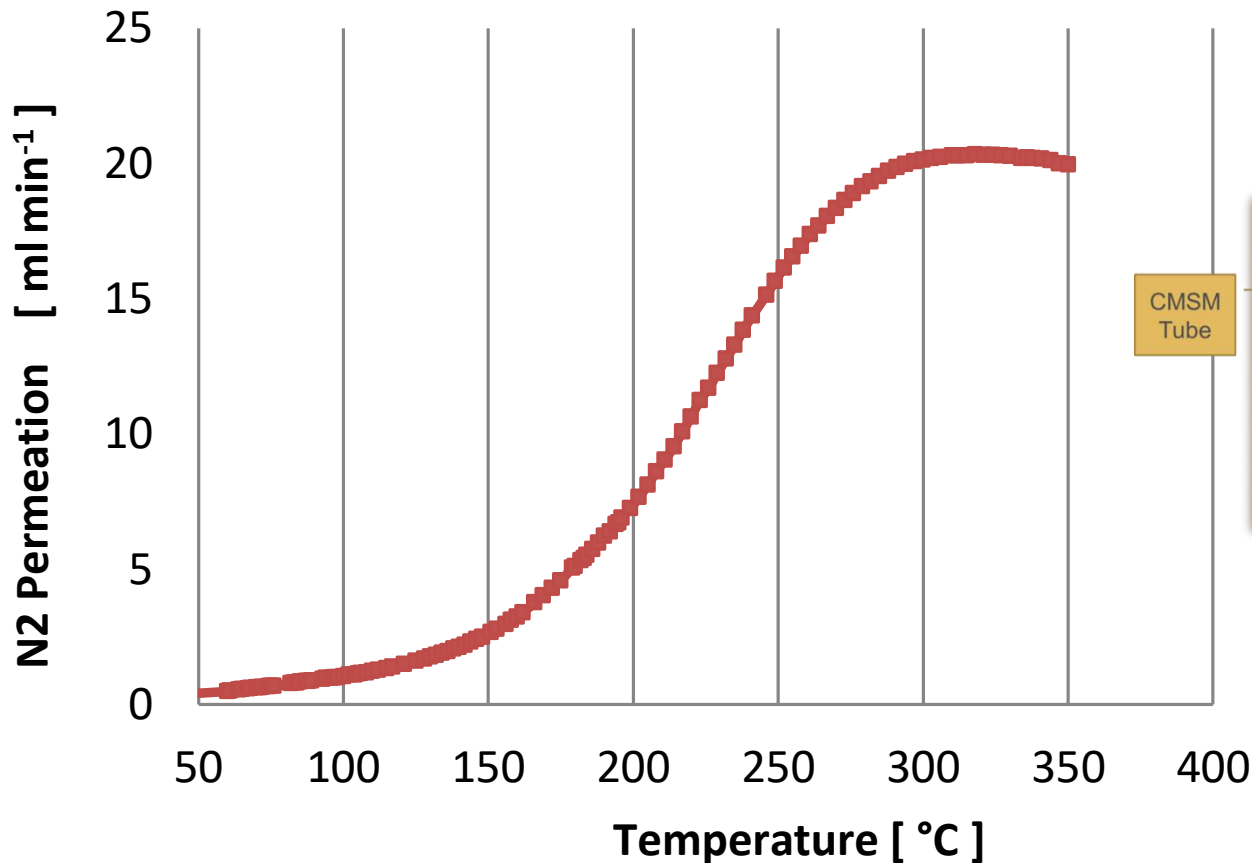
After one day



O₂ and N₂ very low flux

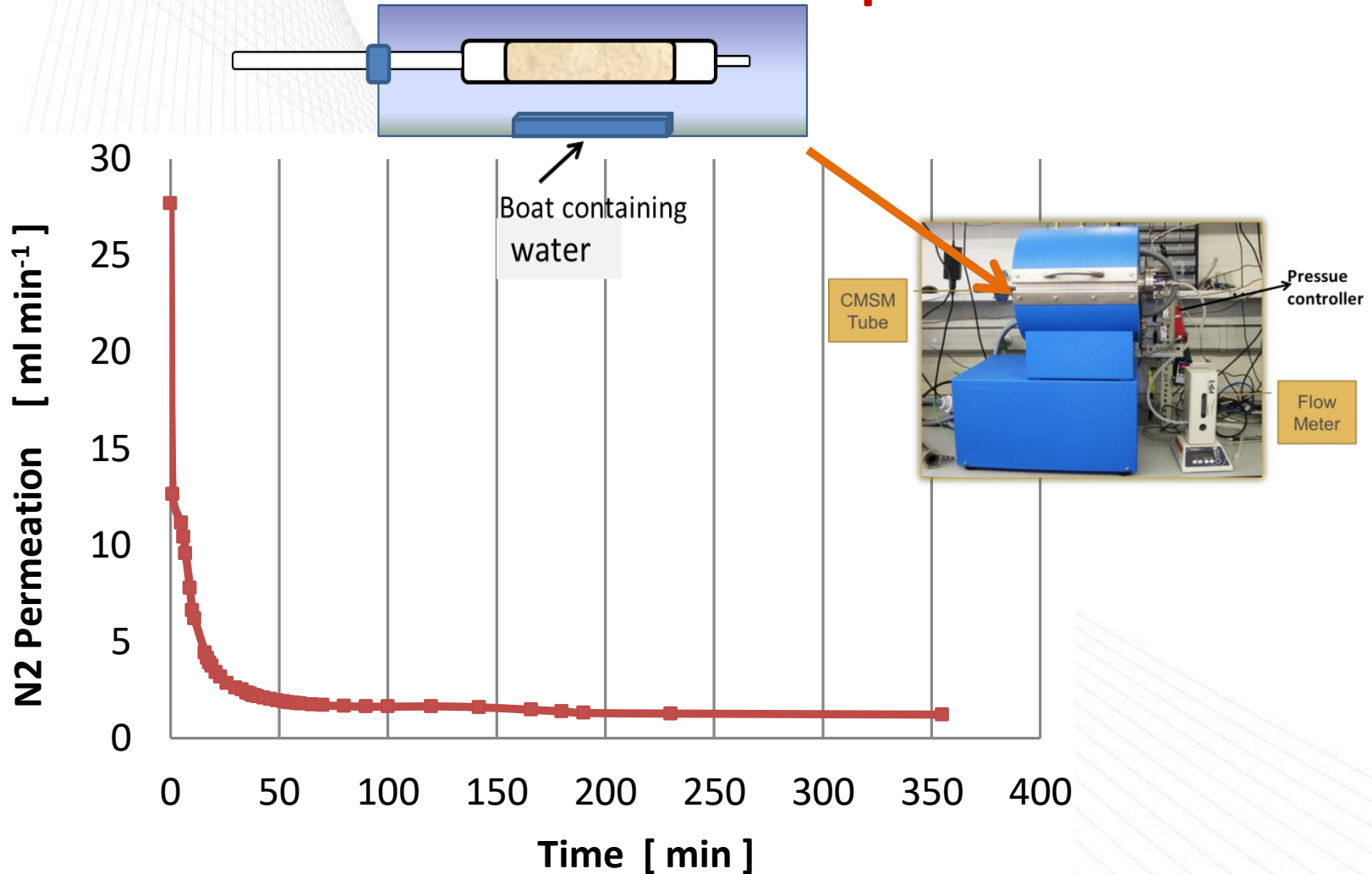
Llosa, Pacheco et.al Int J. hydrogen energy 40 (2015) 5653 40 (2015) 3485

N₂ permeation of c-CMSM-500 in function of the temperature at 400 kpa CMSM-500 after 14 days



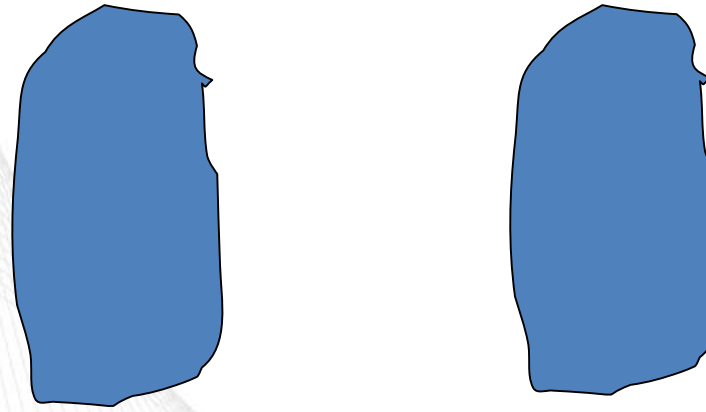
N₂ permeation of fresh membrane 66.8 ml min⁻¹

N₂ permeation at 400 kPa for CMSM-500 in function of the time of water adsorption



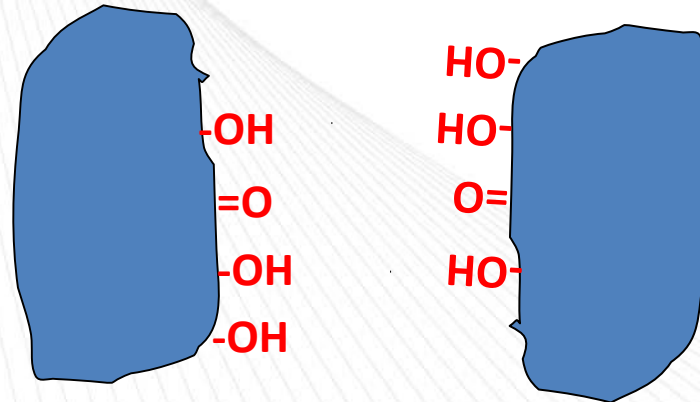
Modification on the porous

Just after
carbonization



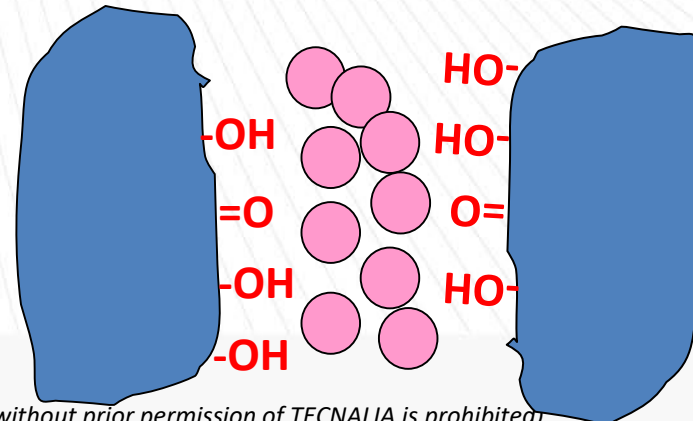
N₂ 66 ml/min

Active places
react with water
(water chemisorption)



N₂ 20 ml/min

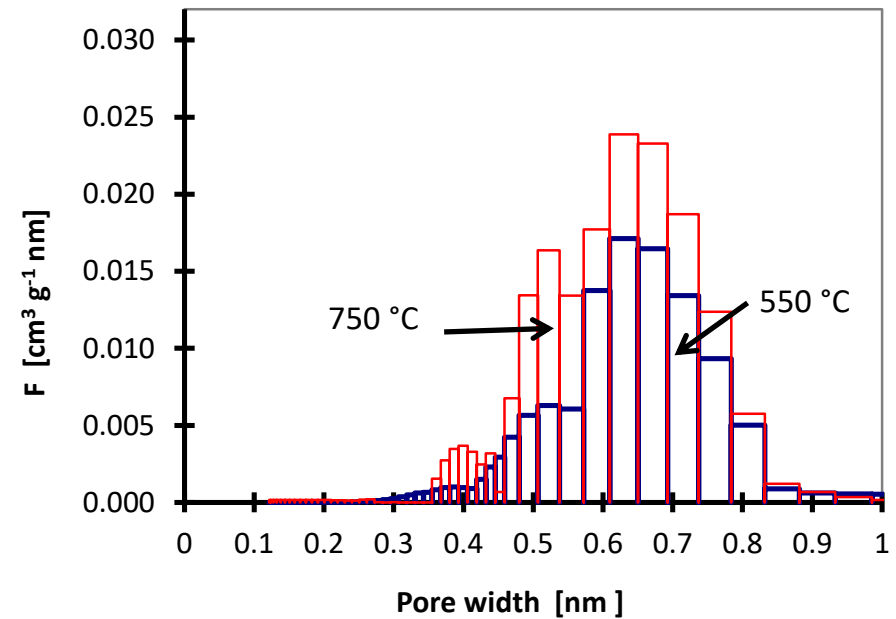
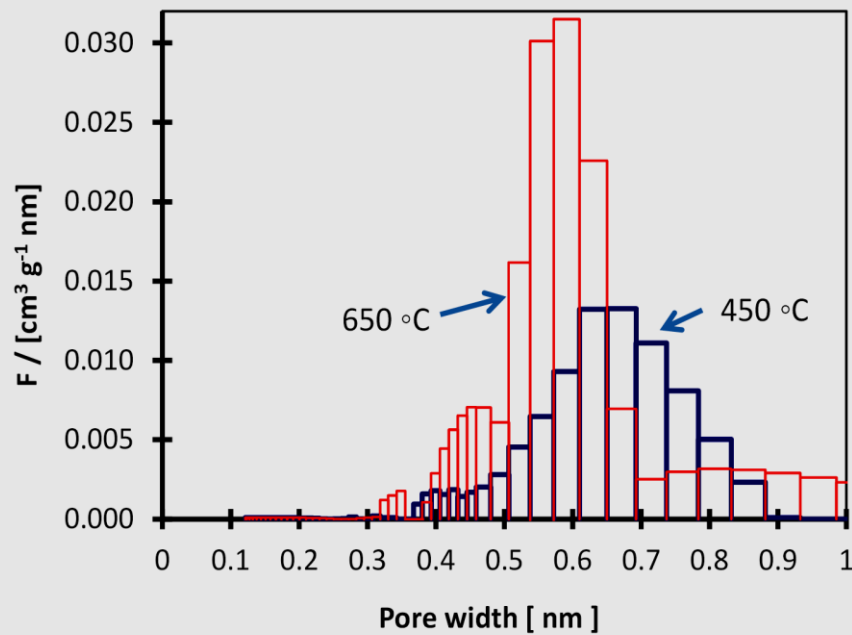
Water physisorption



water

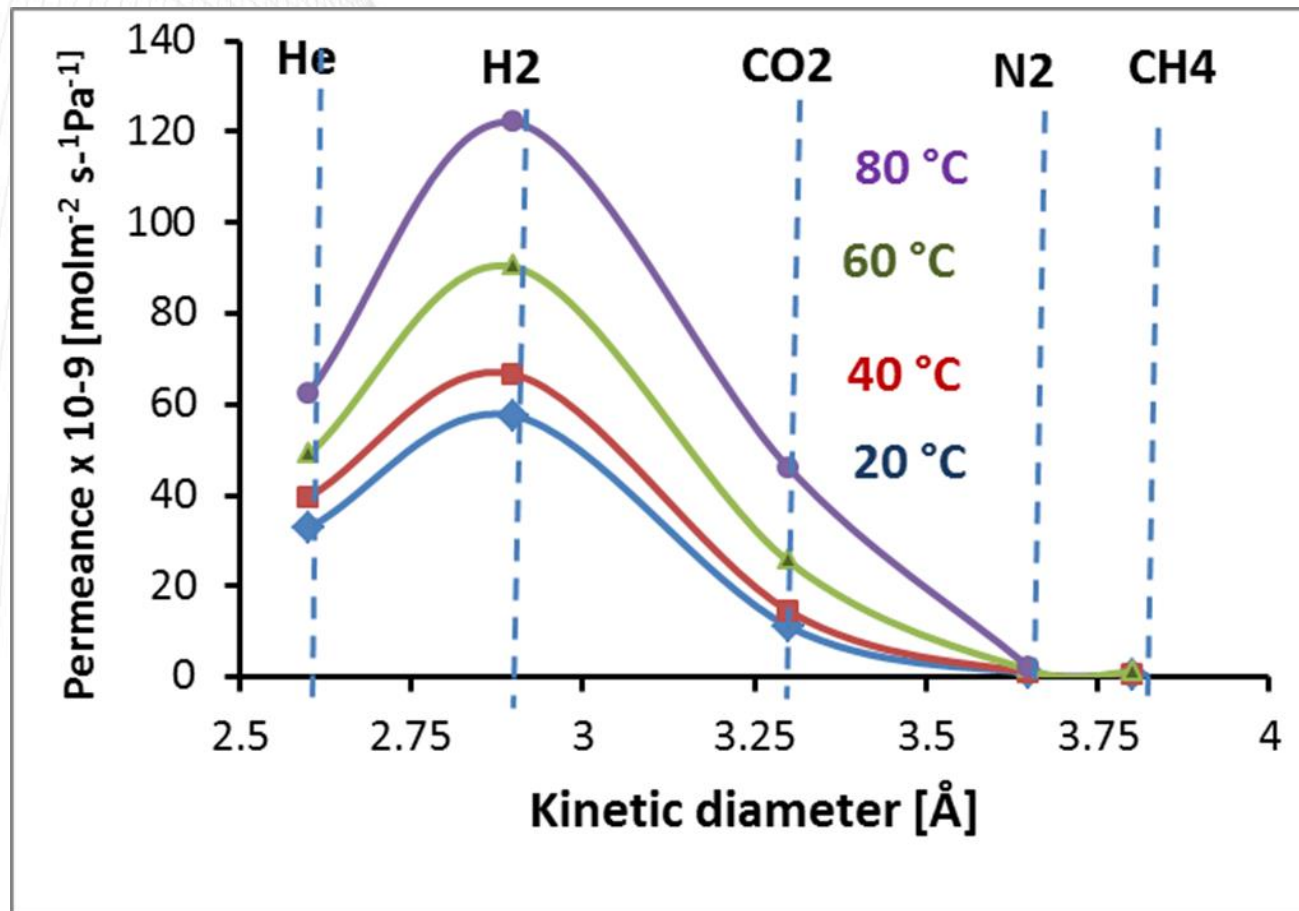
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Pore size distribution at various carbonization temperatures



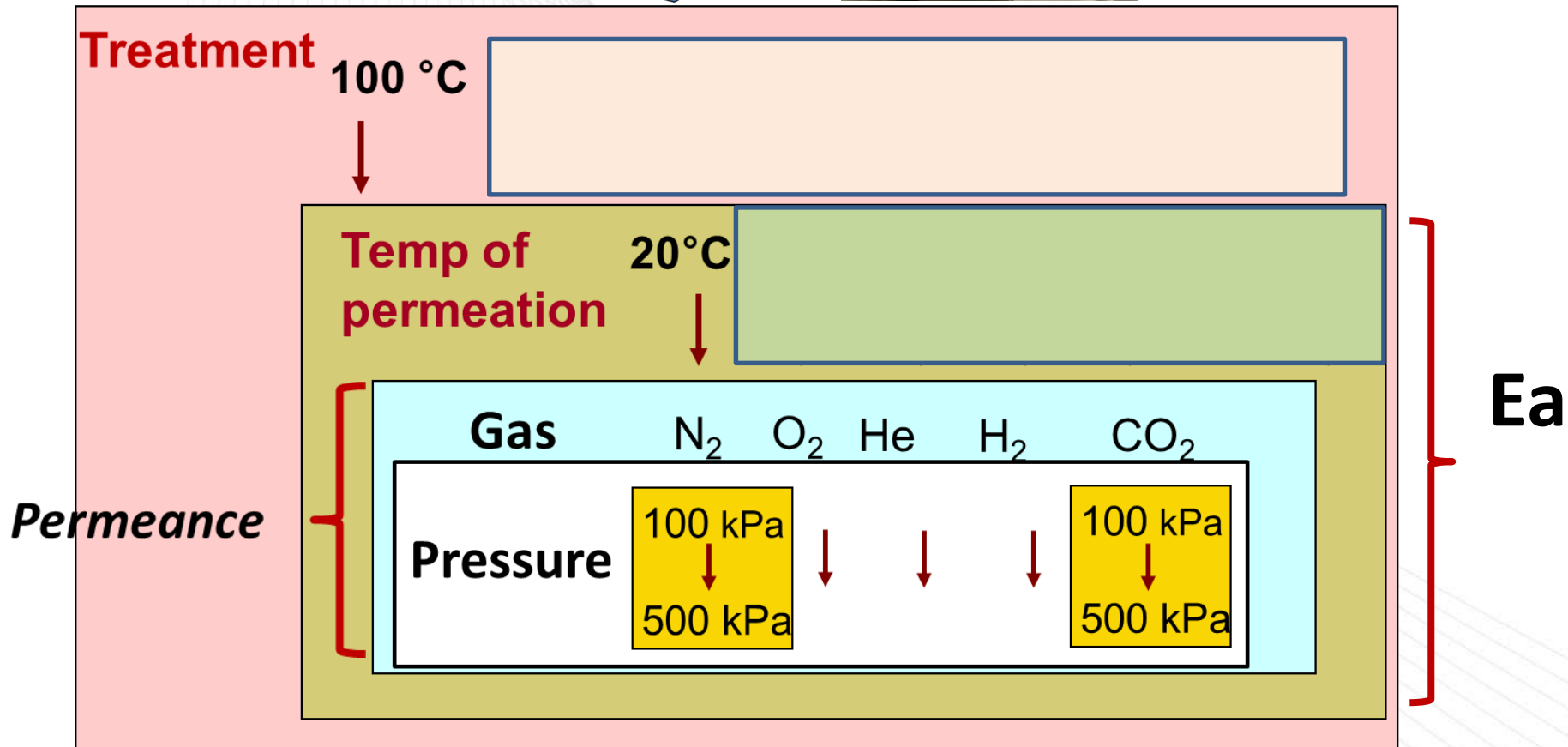
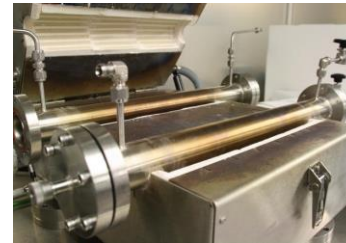
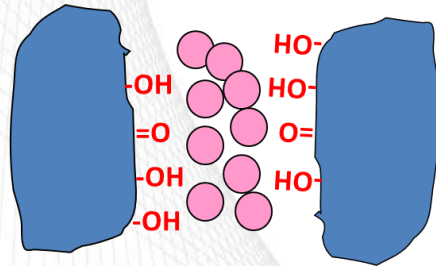
Gas permeance as a function of temperature

Al-CMSM carbonized at 500 °C

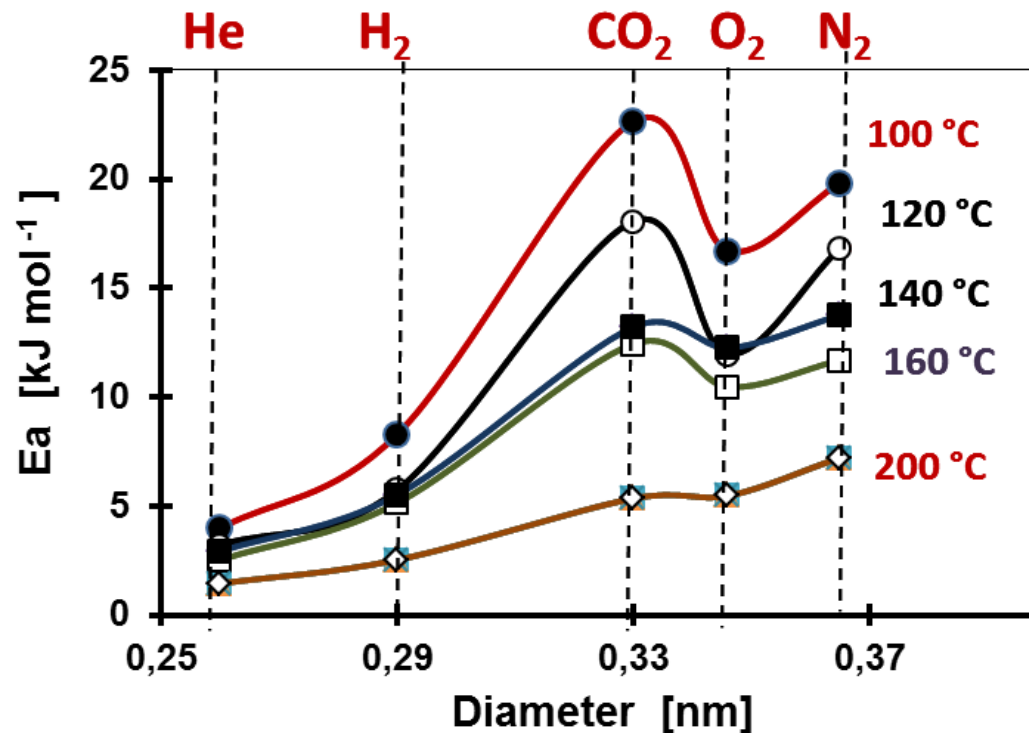
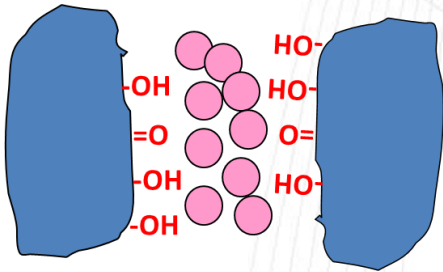


HyGrid

Activation Energy (E_a) at various activation heat treatments



Activation energy of c-CMSM-500 membrane at various temperatures of permeation, for various gases

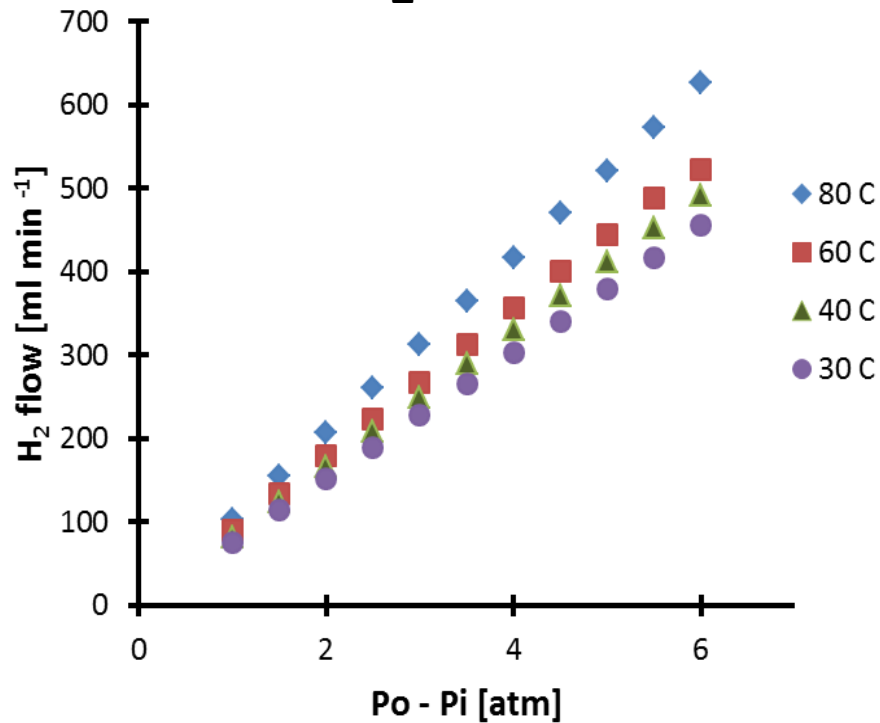


The values correspond to the average of permeation data performed at different temperatures from 40°C to 100°C

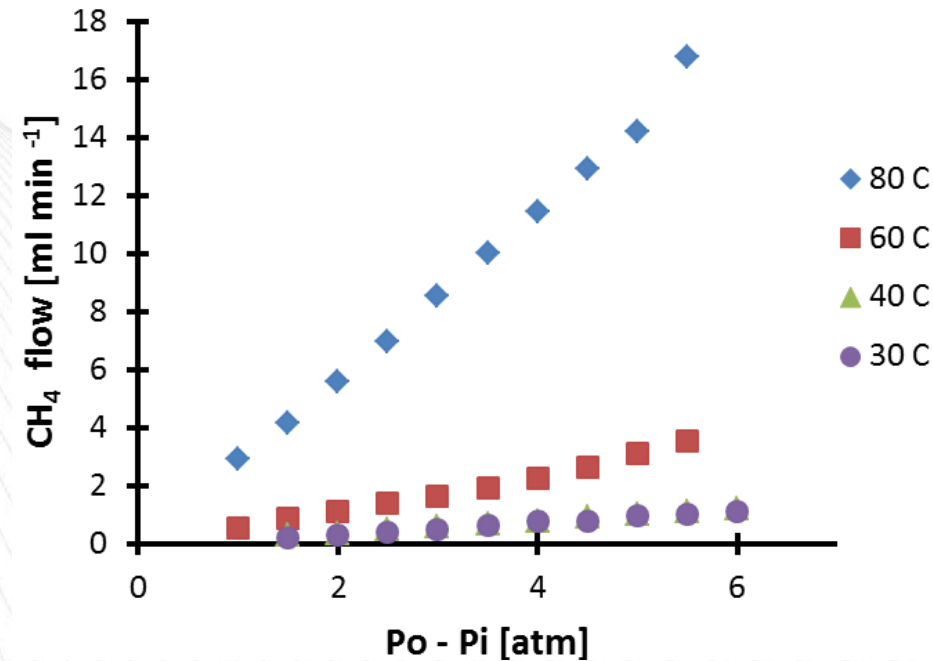
Single H₂ and CH₄ gas flow as a function of temperature and pressure

HyGrid

H₂

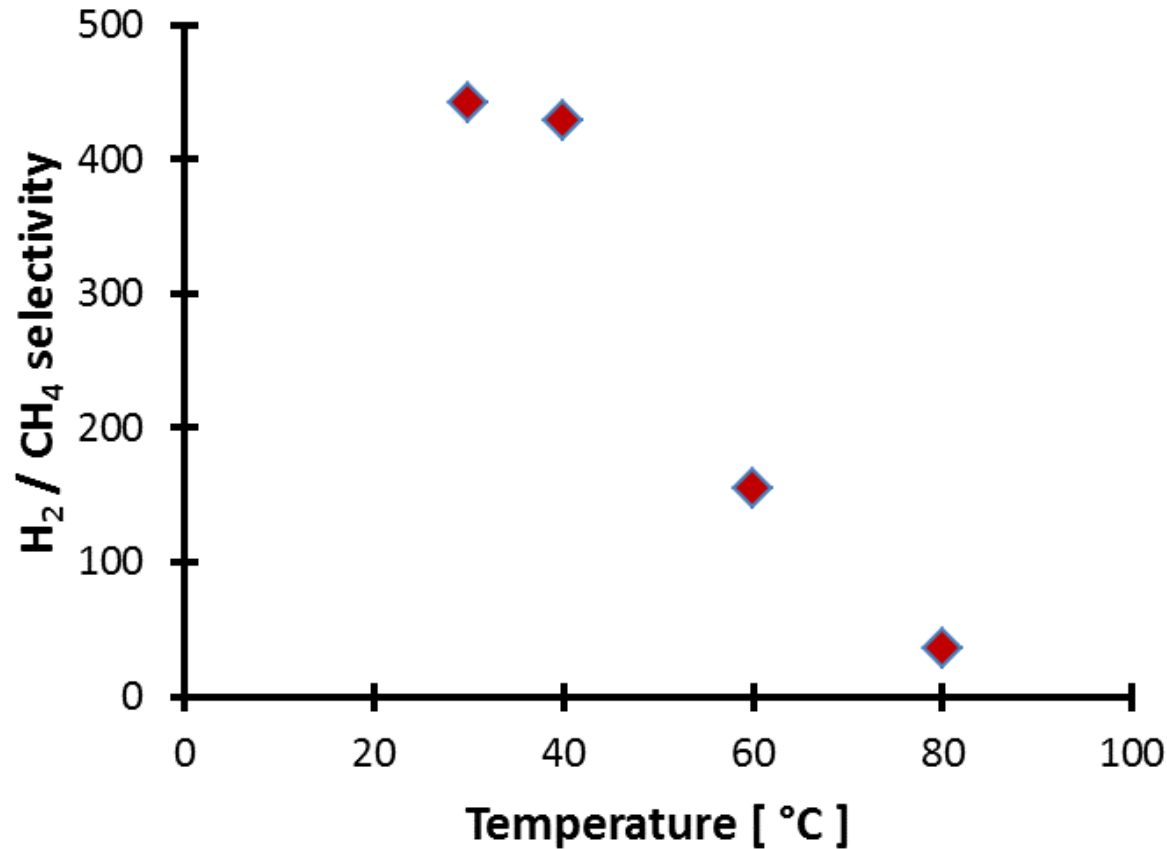


CH₄



H₂ and CH₄ gas ideal perm-selectivity as a function of temperature

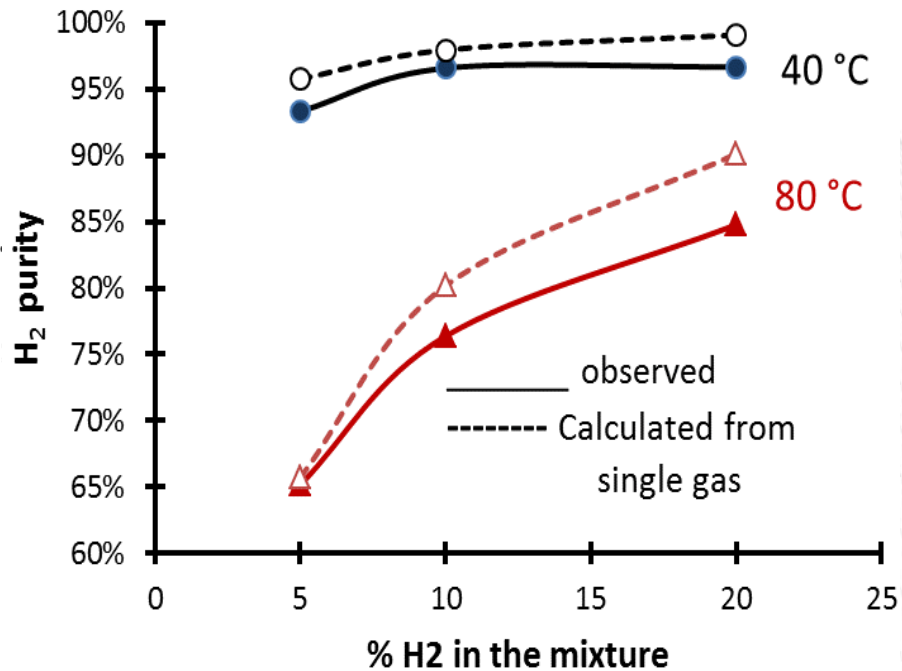
HyGrid



H₂, CH₄ mix gas permeation

H₂ purity (%) at various H₂ contents in the mixture

Pressure: out 7.5, in 0.01 bara, inlet 2 L min⁻¹

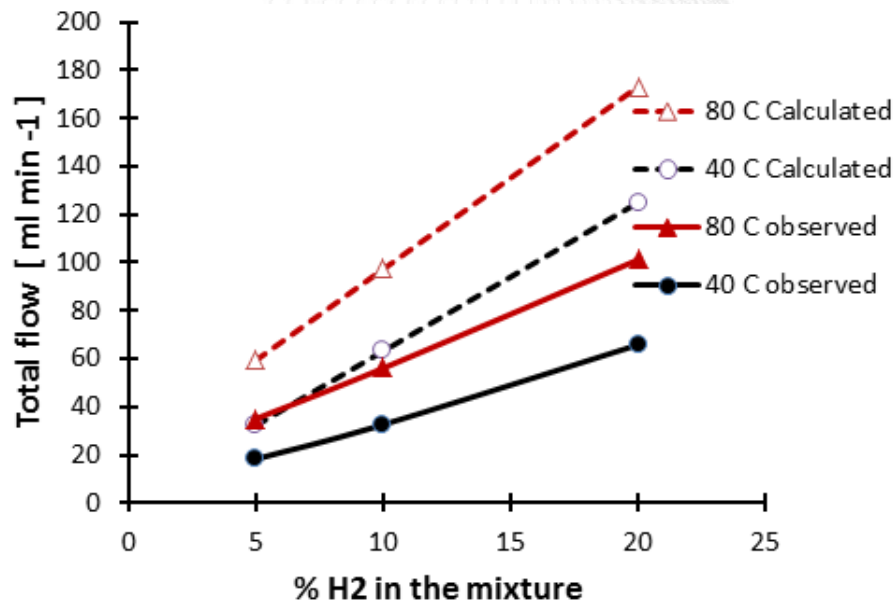


Testing temp (°C)	H ₂ in feed mixture (%)	H ₂ purity in the permeate (%)	
		Measured	Calculated
40	5	93.3	95.7
	10	96.6	98.0
	20	96.7	99.0
80	5	65.2	65.7
	10	76.4	80.2
	20	84.8	90.1

HyGrid

Total flow H₂- CH₄ mixed gas

Pressure: out 7.5, in 0.01 bara, inlet 2 L min⁻¹

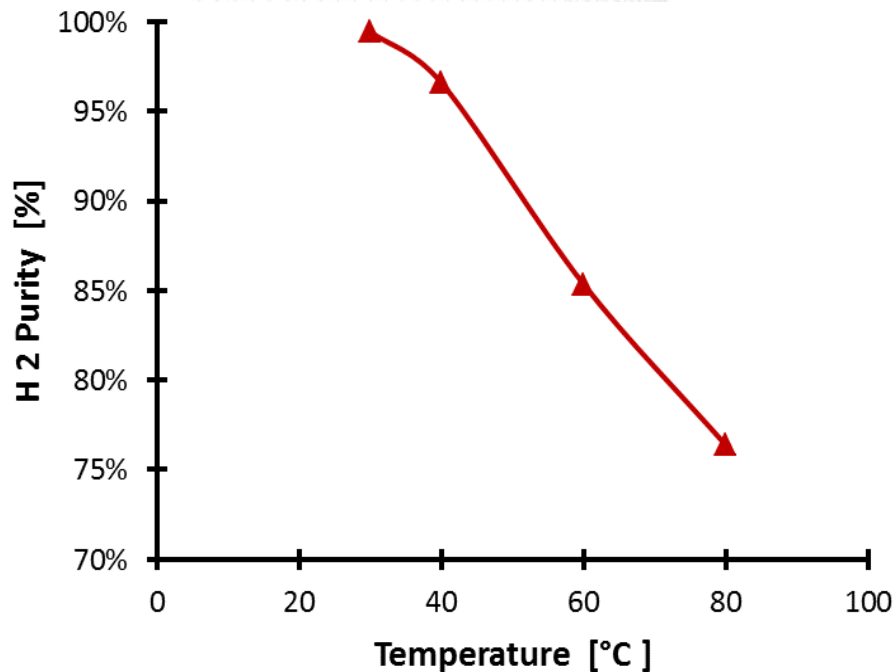


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40	5	93.3	95.7
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	20	96.7	99.0
80	5	65.2	65.7
	10	76.4	80.2
	20	84.8	90.1

HyGrid

H₂-CH₄ mixed gas tests (10% of H₂)

(Pressure_{out}: 7.5 bara; Pressure_{in}: 0.01 bara, inlet 2 L min⁻¹)



Testing temperature (°C)	H ₂ purity in the permeate (%)
80	76.4
60	85.4
40	96.6
30	99.4

Exceeds the H₂ purity requirement of Grade A (98.0 %) and is close to Category 3 (99.9%).

HyGrid



HyGrid



Acknowledgements

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